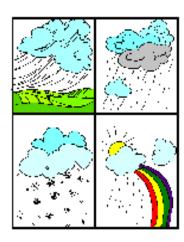
Slide No. 1

Water - a most peculiar substance Describe it

DESCRIBE THE PHASES OF WATER

SOLID



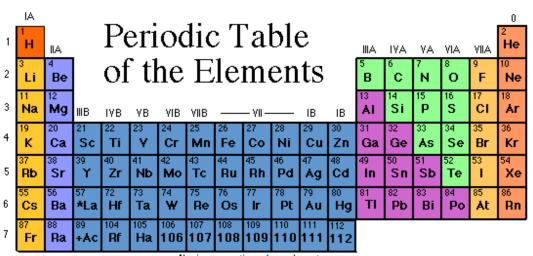
LIQUID

GAS

sublimation evaporation condensation

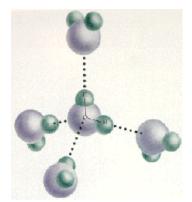
precipitation

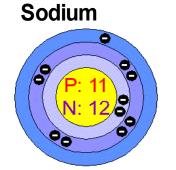
Water is the universal solvent

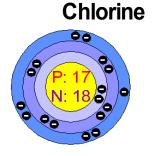


Naming conventions of new elements

*Lanthanide Series														
+ Actinide	90	91	92	93	94	95	96	97	98	99	100	101	102	103
Series	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr







hydrogen bond

ionic bond

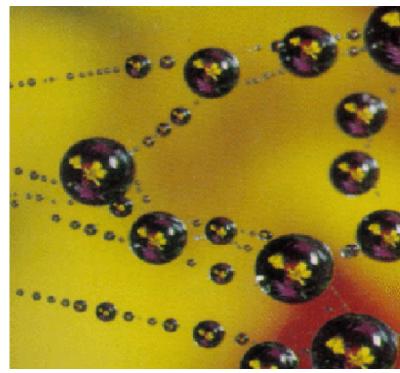
How much of our water supply is

fresh water?

- 1. 1/3
- 2. 3%
- **3.** 10%

How much is ice?

- 1. 2/3
- 2. 1/2
- 3. 1/3

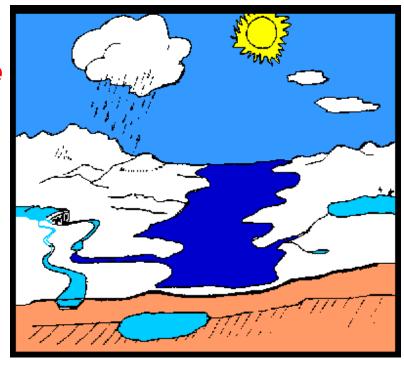


What dissolves in water? Does life depend on it?

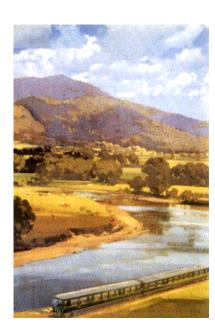
carbon dioxide oxygen

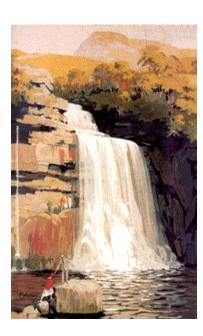
- 1. Water has a high latent heat of vaporization

 Water evaporates slowly from lakes where many
 life forms are dependent on it
- 2. Water is a liquid over a wide temperature range Where does most life live?
- 3. Water is less dense in its solid state than liquid state Life lives under ice
- 4. Water has a very high specific heat (absorb or lose heat before temperature change)



4. Water exhibits viscosity Resistance between layers deeper areas swift current, shallower less Formation of eddies

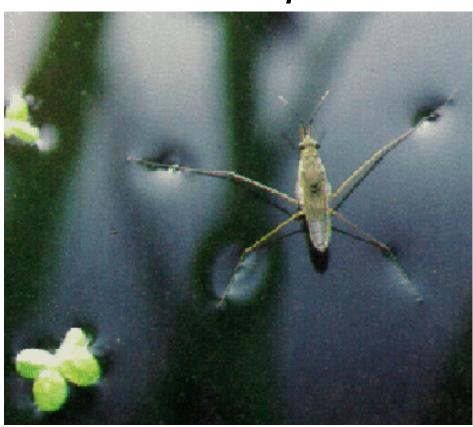




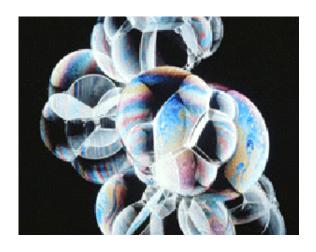


5. Water exhibits surface tension

"Skin" helps organisms to live on the top



SOAP FILM



AIR

ONN
WATER

NOON
WATER

NO

water avoiding
water loving

Soap or Detergent Molecule

H₂O

Slide No. 9

