Periodic Table of the Elements

Alkali Metals — *reactive metals , can explode with water*

Metalloids — *both metals and non metals*

Non Metals — *brittle, no metallic luster, non conductor*

Rare Earth — *many synthetic, 3 electrons in outer*

Halogens — *salt formers, 7 electrons in outer*

Noble — *8 electrons in outer*

P: 8
N: 8
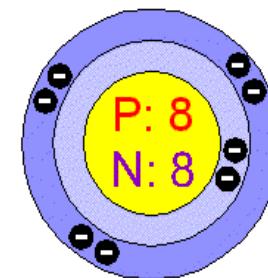
Alkaline — *metallic, very reactive*

Transition Metals — *valence electrons not just in last shell*
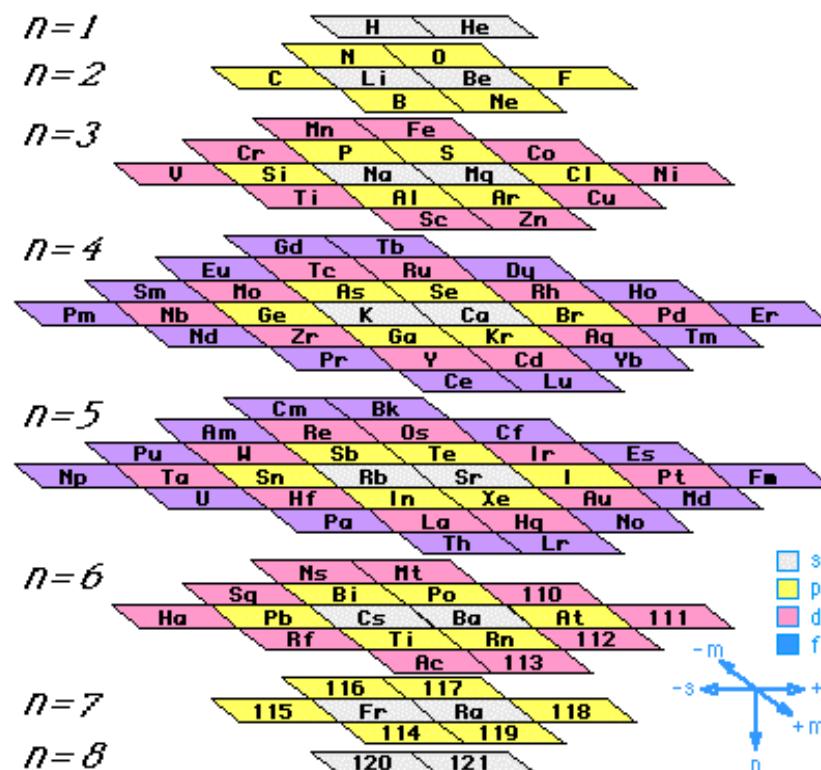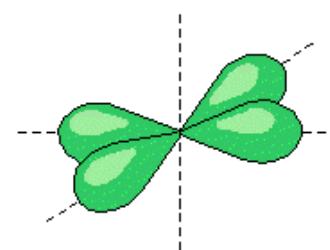
Other Metals — *electrons in outer shell*

How does this chart differ?

Physicist View (Stowe)

Standard Model of Particles

Matter Particles

quarks (up/down; charm/strange: top/bottom)
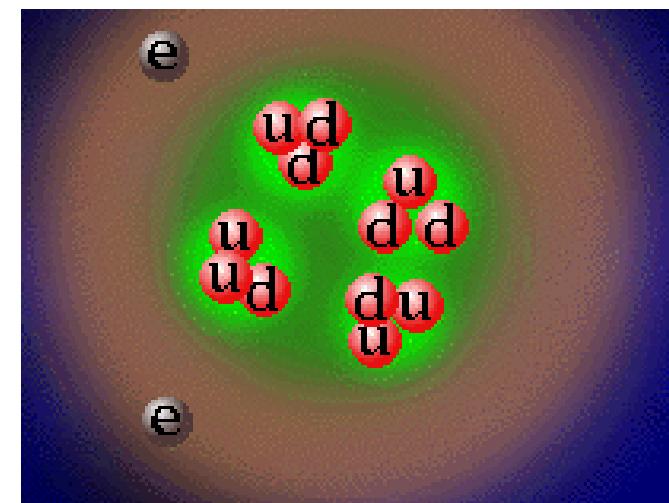
leptons (electron/electron neutrino; muon/

muon neutrino; tau/tau neutrino)

Force Particles
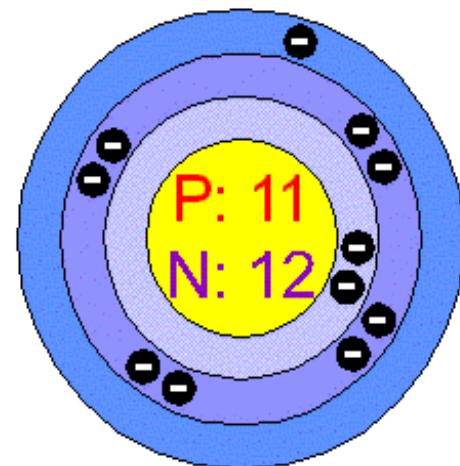
vector bosons (weak)

*gluon* (strong)
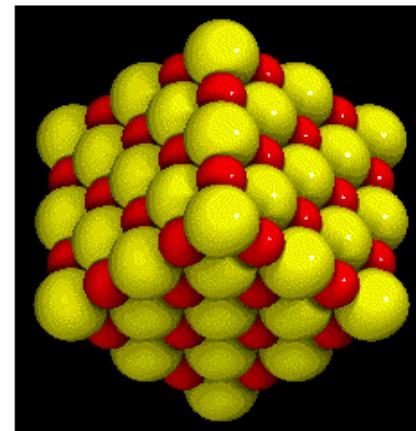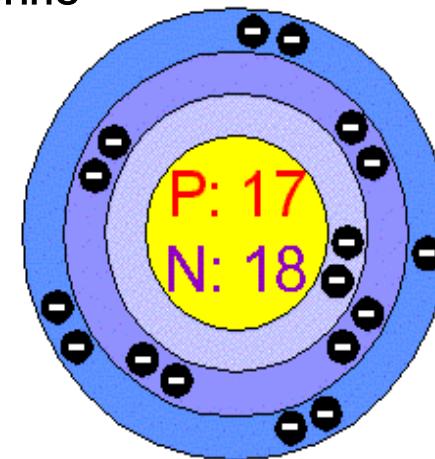
photons (electromagnetic)

Sodium

Chlorine

P: 11
N: 12

P: 17
N: 18

Bohr model
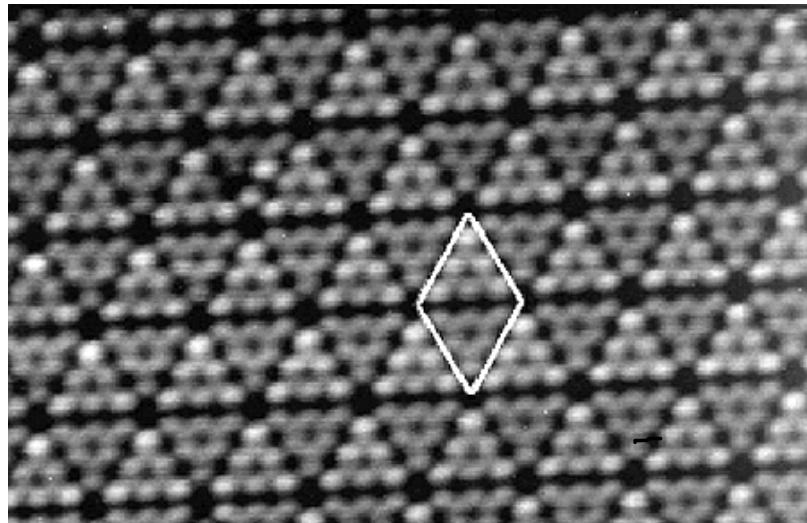
Compound

NaCl (halite)

gold

copper

silicon

Characteristics of elements